



High crystallinity Si-ferrihydrate: An insight into its Néel temperature and size dependence of magnetic properties

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Received 20 June 2006; revised 27 September 2006; accepted 6 October 2006; published 13 February 2007.

[1] Ferrihydrate, an antiferromagnetic iron oxyhydroxide with resulting magnetization due to uncompensated spins, is of great importance for the cycling of many trace metals in the environment. Four ferrihydrate samples prepared with 1.3 to 3.5 wt% of Si at different synthesis temperatures (7.5°C, 22°C, 50°C, and 75°C) were studied by temperature-dependent hysteresis loops, ZFC/FC susceptibility curves, ac susceptibility and Mössbauer spectroscopy. The incorporation of Si into the ferrihydrate during synthesis changed the properties of this mineral. Interestingly, seven sharp lines were observed in the X-ray diffraction pattern of the ferrihydrate samples prepared at 50°C and 75°C.

In general, both XRD and magnetism demonstrate that particle size decreased (from 23 nm to 2 nm) and particle size distribution narrowed as the synthesis temperature was lowered. Those samples prepared between 7.5°C and 50°C showed the expected superparamagnetic behavior of ferrihydrate below 300 K. The ferrihydrate prepared at 75°C was unusually coarse-grained and had a blocking temperature above 300 K. Extrapolation of induced magnetization from the largest particles with the highest crystallinity allowed an estimate of a ferrihydrate Néel temperature of around 422 K. We also present XRD and magnetic data from large natural Si-ferrihydrate collected from a marine shallow-water hydrothermal area that formed at a temperature of approximately 88°C.

Citation: Berquó, T. S., S. K. Banerjee, R. G. Ford, R. L. Penn, and T. Pichler (2007), High crystallinity Si-ferrihydrate: An insight into its Néel temperature and size dependence of magnetic properties, *J. Geophys. Res.*, 112, B02102, doi:10.1029/2006JB004583.

1. Introduction

[2] The study of ferrihydrate is of great interest for its importance in the biogeochemical cycling of iron and associated trace elements and for its apparent role in recording geophysical processes in a range of natural environments. This naturally occurring hydrated ferric oxyhydroxide can be found in a wide variety of different environments such as sediments, mine wastes, acid mine drainage, soils [Jambor and Dutrizac, 1998], Mars surface samples [King and McSween, 2005], and as a precursor of fine-grained magnetic particles in Chinese paleosols [Liu *et al.*, 2005]. It typically occurs as a disordered fine grained nanophase, and it is a common and important constituent of many soils. However, its identification is difficult due to the presence of other minerals and its poor crystallinity. Like other iron oxides it is a strong pigment and can influence many soil properties [Childs, 1992].

[3] In the geochemical cycle, ferrihydrate plays an important role as an adsorber of various trace elements and as a control on the iron cycle in surface waters [Jambor and Dutrizac, 1998]. Rancourt *et al.* [2001] report the occurrence of ferrihydrate containing high concentrations of arsenic. This material forms in shallow water near submarine hydrothermal vents at Tutum Bay, off the coast of Papua New Guinea. This study further suggests that this As-bearing ferrihydrate has a low degree of crystallinity and that coprecipitation of ferrihydrate with As results in retarded particle growth and small particles size. In addition, ferrihydrate is an attractive material for environmental applications due to large surface area, strong adsorptive effects, high adsorption capacity and low cost, and it is often used in processes for treatment and removal of trace elements in wastewaters from various industries [Riveros *et al.*, 2001; Richmond *et al.*, 2004].

[4] Ferrihydrate structure and composition (nominally $\text{Fe}_5\text{HO}_8\text{4H}_2\text{O}$; [Fleischer *et al.*, 1975]) is still a subject of debate even though this antiferromagnetic iron phase has been studied by many authors [e.g., Towe and Bradley, 1967; Drits *et al.*, 1993; Janney *et al.*, 2001; Jansen *et al.*, 2002; Seehra *et al.*, 2004] with different techniques. An initial ferrihydrate precipitate is often produced in the laboratory by the rapid hydrolysis of Fe^{3+} solutions or the slow oxidation of Fe^{2+} solutions. They are usually designated as “2-line” or “6-line” according to the number of characteristic lines observed in their X-ray diffraction

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(XRD) patterns. Those exhibiting six broad peaks are considered more crystalline, whereas those with only two broad peaks are considered poorly crystalline [Cornell and Schwertmann, 2000]. XRD patterns of natural and synthetic ferrihydrite typically exhibit between two and six broad lines.

[5] There is a paucity of reports regarding the fundamental magnetic properties of ferrihydrite, especially with regards to the magnetic anisotropy constant (K) and the Néel temperature (T_N). In the case of the anisotropy constant, reported data for synthetic ferrihydrite range from 35 to 610 kJ/m³ [Suzdalev et al., 1996; Harris et al., 1999; Gilles et al., 2000]. T_N estimates range from 250 to 500 K [Gilles et al., 2002; Jansen et al., 2002; Punnoose et al., 2004; Guyodo et al., 2006] and are, in general, found by extrapolation of saturation magnetization or noncompensated magnetization (the magnetization carried by the uncompensated spins) versus temperature data. Some authors, e.g., Pannalal et al. [2005] claim that ferrihydrite saturates in field of less than 1.2 T and such behavior at room temperature is compatible with a weak ferromagnetic material. Pannalal et al. [2005] further claim that ferrihydrite nanoparticles can act as carriers of Natural Remanent Magnetization.

[6] Ferrihydrite appears to form the core of ferritin, an iron-storage protein of mammals that consists of a hollow, approximately spherical shell of polypeptide with external diameter 12 nm and internal diameter 8 nm [Chasteen and Harrison, 1999; St. Pierre et al., 2001]. The majority of previous work aimed at the magnetic characterization of ferrihydrite has used ferritin [Kilcoyne and Cywinski, 1995; Gider et al., 1995; Tejada et al., 1997; Makhoulf et al., 1997; Friedman et al., 1997; Gilles et al., 2000; St. Pierre et al., 2001; Gilles et al., 2002]. Such particles can be considered as an ensemble of magnetically noninteracting particles [Gilles et al., 2000]. In contrast, synthetic samples of ferrihydrite are commonly aggregated, which means that the particles are probably magnetically interacting. This complicates the magnetic characterization of these samples because magnetic dipolar interactions are likely present.

[7] The ferrihydrite found in nature can vary considerably in composition and often contains substantial amounts of coprecipitated silicon. The presence of silicon-bearing ferrihydrite (Si-Fh) in soils is documented by Childs [1992], who has reviewed the available information of XRD and Mössbauer spectroscopy of ferrihydrite. Special attention is also paid to the association between ferrihydrite and Si, and Childs [1992] suggests that natural ferrihydrites from soil-related environments commonly contain up to 9% Si. There are still some doubts whether presence of Si is structural or if it is adsorbed on the surface in natural material [Parfitt et al., 1992]. Anderson and Benjanim [1985] and Campbell et al. [2002] demonstrated the ability of Si to inhibit ferrihydrite transformation to other more stable iron phases like goethite and hematite. Also, Campbell et al. [2002] observed the formation of silicon-bearing hematite as a result of heating Si-containing ferrihydrite at 800°C. The observation by Campbell et al. [2002] that Si was structurally incorporated into hematite produced by thermal transformation of Si-Fh lead them to conclude that Si was also structurally incorporated in the precursor ferrihydrite. In contrast, synthetic samples of Si-ferrihydrite were

reported by Childs [1992] and Schwertmann et al. [2004] to present an unusual XRD pattern: seven well-defined peaks that are significantly sharper, indicating a greater degree of crystallinity.

[8] Here we present results of the magnetic characterization of four synthetic ferrihydrite samples prepared with 1.3 to 3.5 wt % Si by slow oxidation of dissolved Fe²⁺ at different temperatures (7.5°C, 22°C, 50°C, and 75°C) and one natural ferrihydrite sample collected from a marine shallow-water hydrothermal area. XRD patterns of ferrihydrite samples synthesized at temperatures at and above 50°C as well as the natural hydrothermal ferrihydrite showed seven sharp lines. These samples have structural features similar to those of the samples described by the authors above. Characteristic magnetic blocking temperatures and distributions of particle size were observed for the 7.5°C, 22°C, 50°C, 75°C and the natural ferrihydrite samples by using low temperature magnetic measurements and ac susceptibility. Additional information was obtained by Mössbauer spectroscopy, which, combined with the methods above, has provided useful information about particle size, size distribution, and inferences regarding Si structural incorporation in ferrihydrite. The Néel temperature for ferrihydrite was also estimated from magnetic measurements performed on the 75°C sample.

2. Material

2.1. Synthetic Ferrihydrite

[9] Silica-bearing ferrihydrite samples were prepared by air oxidation of ferrous iron dissolved in an aqueous medium containing dissolved silica. First, the aqueous medium was prepared by adding the following reagent grade components to one liter of deionized water: 0.0142 g KCl, 0.2084 g CaCl₂, 0.0618 g MgSO₄, 0.0708 g NH₄HCO₃, 0.1680 g NaHCO₃, 0.1767 g Na₂SO₄, 0.2188 g 2 N HCl, 0.626 g 0.16 N H₂SO₄, and 10 ml of a 1000 μg Si/ml aqueous silica standard (sodium silicate in water). The solutions were allowed to mix using a magnetic stirrer with Teflon coated stir bar until the salts dissolved, and the resultant solution was filtered under vacuum through a 0.2 μm Millipore membrane filter to remove precipitated calcite. A pH-buffered aliquot of this aqueous medium was then prepared by adding 3.9066 g of the weak acid 2-(4-morpholino) ethane sulfonic acid (MES) to 200 mL of the filtered synthetic medium and adjusting the pH to 6 (pK_{a,MES} ~ 6.1 at 25°C) through addition of 1 M NaOH (typically 8 mL). The temperature of the synthetic medium was equilibrated in a closed vessel at the desired reaction temperature prior to the addition of ferrous iron (7.5°C in refrigerator, 22°C (ambiente), and 50°C/75°C in water bath). Ferrihydrite precipitation was initiated in the presence of Si (356 μmol Si/L) and the other matrix components through direct addition of solid ferrous chloride salt (0.0351 g of FeCl₂·4H₂O; 1165 μmol Fe/L) and mixing was achieved using an overhead paddle stirrer (~500 rpm). Thus the synthesis experiments were conducted with a starting Si:Fe ratio in solution of 0.31. Evaporation was insignificant at elevated temperatures due to the short synthesis time. After approximately 5 min, a 1 mL sample of the experimental solution was collected, filtered (0.2 μm nylon syringe filter) and the concentration of dissolved Fe²⁺ was measured colorimetrically using

1,10-phenanthroline at 510 nm on a Hach DR/2010 spectrophotometer. Ferrous iron measurements were repeated at regular intervals until at least 50% of the ferrous iron (introduced as ferrous chloride) had been oxidized. At the end of the oxidation-precipitation experiment, the iron oxide solids were collected by vacuum filtration on a 0.2 μm polycarbonate membrane filter. The filtered solids were immediately washed with deionized water and then allowed to air-dry. The dried solids were scraped from the filter surface and stored for further characterization. Multiple syntheses were conducted at each reaction temperature to ensure reproducibility, which was confirmed through calculation of Fe^{2+} oxidation rate along with mineralogical characterization.

[10] For simplicity, Si-Fh samples are denoted as *Fh7.5*, *Fh22*, *Fh50*, and *Fh75*, with each number serving as a reference for each synthesis temperature (7.5°C, 22°C, 50°C, and 75°C, respectively).

2.2. Natural Ferrihydrite

[11] The natural ferrihydrite sample, *BC-11*, was collected in the spring of 2003 at a marine shallow-water hydrothermal system in Bahia Concepcion, Baja California, Mexico. The shallow-water hydrothermal system consists of several small vents approximately 30 to 50 m offshore at a depth of 7 to 10 m. The hydrothermal fluid pH and temperature were measured at three different vents and ranged from 5.9 to 6.1 and 87 to 89°C, respectively. The vent fluid salinity and chemical composition were similar to those of seawater, with the exception of Si, Fe, and As. These three elements were enriched relative to seawater by up to 500 times, a common phenomenon in marine shallow-water hydrothermal systems [e.g., *Pichler et al.*, 1999a].

[12] Sample *BC-11* is part of the hydrous ferric oxide (HFO) deposits that formed around the vent sites. Those deposits were present throughout the area of hydrothermal activity where they formed as very thin layers on sediment grains in areas of high seafloor temperature. Massive layers and extensive filling of sediment pore space, however, were restricted to the vicinity of vent sites. Deposits varied in color from a bright orange to very dark brown that is almost black, and in Moh's hardness from <1 to about 2.5, similar to the marine shallow-water hydrothermal system in Tutum Bay, Papua New Guinea [*Pichler et al.*, 1999b]. Chemically, *BC-11* also compares well to ferrihydrite deposits in Tutum Bay with Fe (~40 wt%), Si (~10 wt%), and As (~5 wt%) being the most abundant elements.

3. Experimental Details

[13] A commercial SQUID magnetometer (MPMS-XL; Quantum Design) was employed to perform static (ZFC/FC susceptibility curves, hysteresis loops) and dynamic (ac susceptibility curves) measurements as a function of field, temperature and driving frequency. Zero-field-cooled (ZFC) magnetization curves were obtained by cooling in zero field from a high temperature (300 K) to a low temperature (2 K) and then measuring the magnetization at stepwise increasing temperatures from 2 K to 300 K in a small applied field ($B = 5$ mT). The sample was again cooled, in the same field, and field cooled (FC) magnetization curves were obtained by measuring at stepwise increasing temperatures, from 2 K

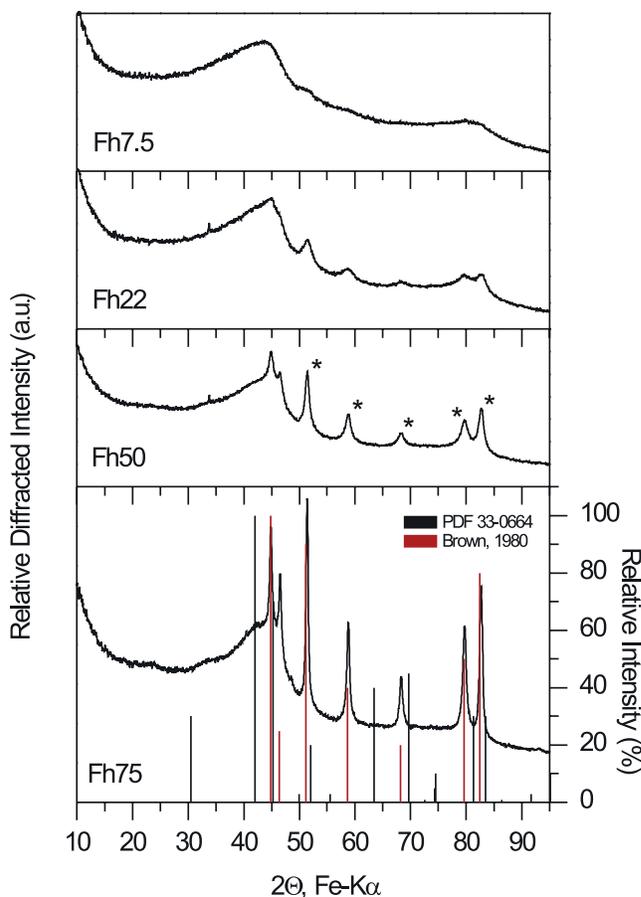


Figure 1a. XRD for all four samples showing the seven-line patterns reflecting particle size increase for samples prepared at different temperatures. Relative peak intensities for hematite (PDF 33-0664) and crystalline ferrihydrite [Brown, 1980] are included for reference. The stars indicate peaks used to calculate crystallite dimensions based on the Scherrer formula; only four peaks could be fit for sample *Fh7.5*.

to 300 K. Hysteresis loops were obtained by using maximum applied fields up to 5 T at temperatures of 2, 100, 200, and 300 K. The ac susceptibility (χ') curves were measured in the temperature range 2–400 K. Mössbauer spectra were measured at room temperature and at 4.2 K. A conventional constant-acceleration spectrometer was used in transmission geometry with a $^{57}\text{Co}/\text{Rh}$ source, using $\alpha\text{-Fe}$ at room temperature to calibrate isomer shifts and velocity scale.

[14] Dried sample powders were characterized by X-ray diffraction using a Rigaku Miniflex diffractometer. Dried solids (10–15 mg) were transferred to a zero-background quartz slide. The solids were then saturated with methanol, spread evenly on the surface of the slide, and then allowed to air-dry. Diffraction data were collected using Fe $\text{K}\alpha$ radiation (30 kV, 15 mA) over an angular range of 5–95° 2θ with a 0.02° step size (12 s/step). Goniometer alignment was checked externally using NIST SRM640b (silicon) prepared in the same manner.

[15] Samples for high-resolution transmission electron microscopy (HRTEM) were prepared by two methods. First,

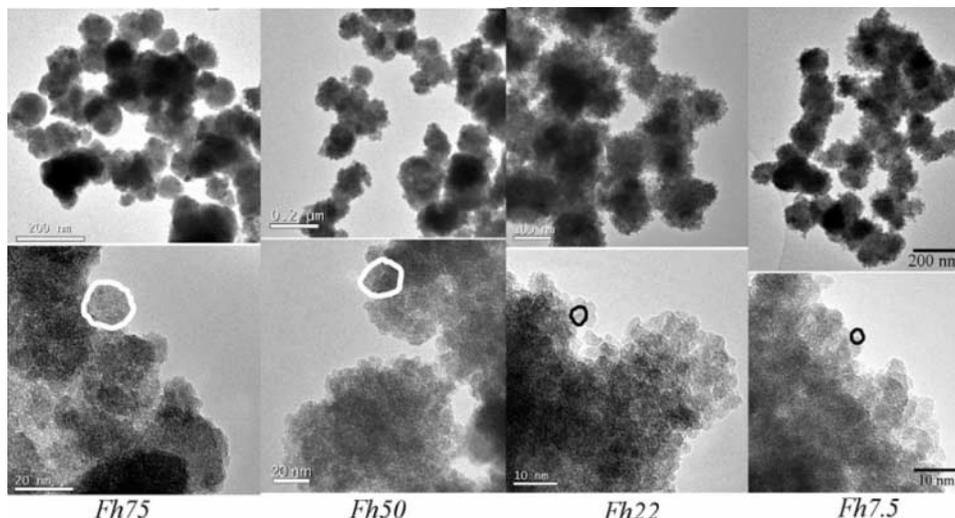


Figure 1b. HRTEM images show the samples are constituted of aggregated (upper pictures) particles and there is a trend of decrease in the primary particle size (circles) with decrease of synthesis temperature.

the dried samples were simply resuspended in purified water, vortexed for a few minutes, and sonicated for a few minutes. While the powders did seem to resuspend readily, a second preparation method was adopted in order to attempt to compare the aggregation state of the particles before and after the drying step. The second method involved syringe filtering the ferrihydrite suspension through a 0.2 μm Nuclepore polycarbonate track etch filter membrane. The supernatant was clear and colorless. Filtration was followed by passing purified water through the filter in order to remove dissolved side-products, salts, and buffers from the particles, and the liquid passed through the filter was again clear and colorless. The filter assembly was disassembled, and the filter was placed in 1 mL of purified water and sonicated for several minutes and vortexed for several minutes. The particles were readily resuspended. A small amount of the suspension resulting from each method was diluted and a single drop placed onto a 200 mesh holey carbon-coated copper grid (SPI), which was then allowed to dry in air. TEM images were collected using either an FEI Tecnai T12 TEM operated at 120 kV or an FEI Tecnai G2 30 operated at 300 kV. All images were collected using a charge-couple device (CCD) camera, and high-resolution images were collected using the lowest intensity conditions possible.

4. X-Ray Diffraction and TEM

[16] Figure 1a shows the XRD patterns obtained for synthetic samples precipitated in the presence of silica. The X-ray diffraction (XRD) patterns for Si-bearing samples synthesized at 7.5°C and 22°C are represented by the usual broad lines characteristic for ferrihydrite found in soils and natural environments [e.g., Carlson and Schwertmann, 1981; Schwertmann et al., 1982]. For the other two samples prepared at 50°C and 75°C, unusual diffraction patterns with seven, comparatively sharp lines were obtained. The XRD pattern for the *BC-11* ferrihydrite sample also showed seven lines (Figure 2a) and the sharp lines suggest the

presence of a well-crystallized material. These patterns are similar to that presented by Childs [1992], which exhibited seven sharp lines. Comparison of our diffraction data for samples *Fh75* and *BC-11* to peaks identified by Brown [1980] reveals consistency both in peak location and relative intensity. Results shown in Figure 1a and documented in Table 1 demonstrate that diffraction peaks identified in our sample patterns can be attributed solely to ferrihydrite. The sharp peak identified at ~ 2.54 Å could potentially be associated with goethite (2.45 Å indexed as (111) in PDF 29-0713), but no other peak is present within the pattern to confirm the presence of this phase (e.g., 4.183 Å indexed as (110) in PDF 29-0713). It should be noted that the peak at ~ 2.54 Å has been indexed within recently proposed structural models for ferrihydrite [Drits et al., 1993; Janney et al., 2001]. In addition, review of published diffraction data for synthetic and natural “6-line” ferrihydrites consistently reveal the presence of a shoulder on the high two-theta side

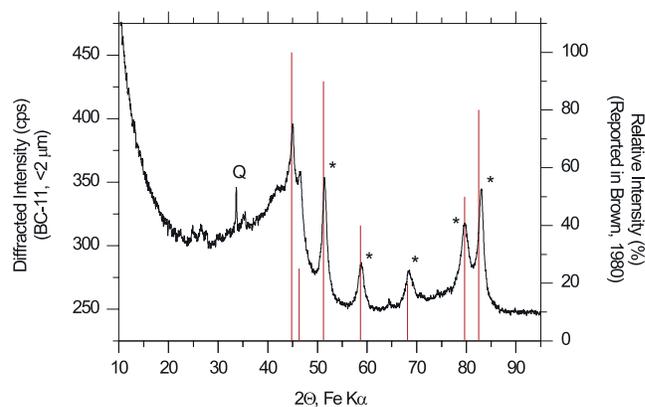


Figure 2a. Characterization of a natural “7-line” ferrihydrite collected from a marine shallow-water hydrothermal area near Bahia Concepcion. The stars indicate peaks used to calculate crystallite dimensions based on the Scherrer formula; Q refers to quartz.

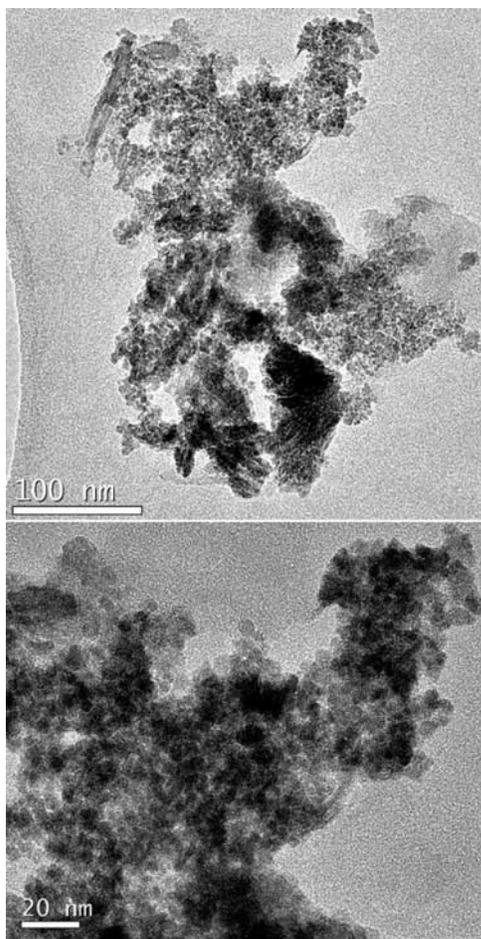


Figure 2b. HRTEM images for natural ferrihydrite (sample *BC-11*).

of the main ferrihydrite peak located at ~ 2.45 Å. Our diffraction data confirm that ferrihydrite is the only detectable phase within our synthetic samples. In samples *Fh50* and *Fh75*, there is also a notable broad reflection located in the region of 40–50 degrees two-theta (Fe $K\alpha$ radiation). This peak is similar to that observed in the diffraction pattern for the poorly crystalline ferrihydrite (sample *Fh7.5*). It thus appears that samples *Fh50* and *Fh75* consist of a heterogeneous mixture of material consistent with larger and more crystalline seven-line ferrihydrite and a less crystalline material with much smaller particle size. This is consistent with the notion of ferrihydrite particle growth proposed by *Parfitt et al.* [1992] in the presence of silica, i.e., our samples *Fh50* and *Fh75* represent mixtures of ferrihydrite particles of varying size similar to the schematic models presented by these authors (their Figure 4). As discussed below, the magnetics data collected for these samples supports this observation.

[17] In general, HRTEM characterization shows the synthetic and natural ferrihydrite particles are heavily agglomerated (Figures 1b and 2b). While it is difficult to generate a particle size analysis from these images, a general trend of increasing primary particle size (a few to several nm in diameter and represented by circles in Figure 1b) with increasing temperature is evident while the size of the

agglomerates seems independent of temperature. Lattice fringes in the HRTEM images confirm that these materials are crystalline and confirm that the single crystal size increases with increasing temperature. Interestingly, single crystal domains appear to be composed of an increasing number of primary particles with increasing oxidation temperature, as evidenced by features like dimples and the incorporation of defects. This could be the result of an increase in oriented aggregation or solid state recrystallization at the higher temperatures. As features like dimples, pores, and slight misorientation between crystalline primary particles are retained, the former is probably more likely [Penn, 2004]. HRTEM characterization for sample *BC-11* (Figure 2b) shows the particles are not well dispersed.

5. Magnetic Study

5.1. ZFC/FC Susceptibility and Hysteresis Loops

[18] Figure 3 shows ZFC/FC susceptibility curves from all four synthetic samples, and three of the four samples exhibit typical superparamagnetic behavior (e.g., as discussed by

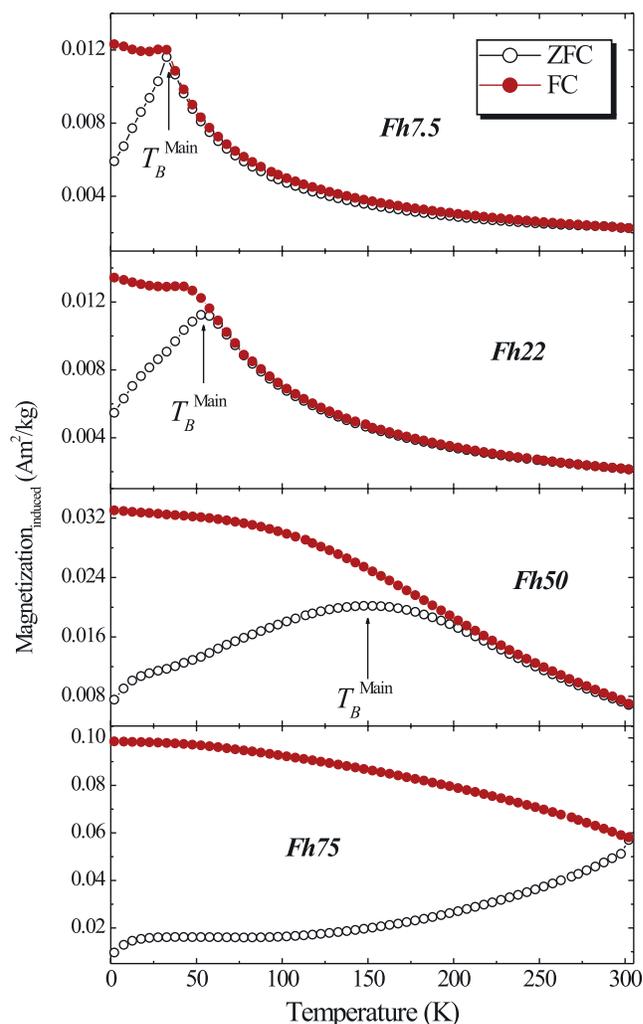


Figure 3. ZFC/FC curves with applied field of 5 mT showing the decrease of T_B^{Main} with the synthesis temperature. Note that T_B^{Main} for *Fh75* is clearly greater than 300 K.

Table 1. Reported d-Spacings (Å) for Peaks Identified in Powder Diffraction Patterns for Ferrihydrite and Other Iron Oxides With Related Structures

Ferrihydrite Samples						ICDD Powder Diffraction File				
<i>Fh75</i> (Fe Kα) ^a	<i>BC-11</i> (Fe Kα) ^a	<i>Childs</i> [1992] (Co Kα) (FH9037) ^b	<i>Brown</i> [1980]	<i>Stanjek and Weidler</i> [1992] (Co Kα)	<i>Janney et al.</i> [2000] (Cr Kα)	Ferrihydrite PDF 29-712	Feroxyhyte ^c PDF 13-0087	Schwertmannite ^d	Hematite PDF 33-0664	
(4.76)	–	4.6	–	4.55	4.56	–	4.61	4.86	–	
(3.40)	–	3.4	–	3.32	3.21	–	–	3.39	3.684	
–	–	–	–	2.717 ^e	–	–	–	–	2.7	
2.541	2.523	2.54	2.54	2.517	2.52 ^f	2.50	2.545	2.55	2.519	
2.454	2.446	2.46	2.46	–	[<i>Janney et al.</i> , 2001] ^g	–	–	–	–	
2.236	2.231	2.28	2.24	2.239	2.22	2.21	2.255	2.28	2.292	
–	–	–	–	–	–	–	–	–	2.207	
1.974	1.966	1.98	1.977	1.977	1.96	1.96	–	1.95	2.077	
–	–	–	–	–	–	–	–	–	1.84	
1.726	1.721	1.73	1.727	1.699	1.70	1.72	1.685	1.66	1.694	
–	–	–	–	–	–	–	–	–	1.636	
–	–	–	–	–	–	–	–	–	1.603	
1.511	1.509	1.51	1.512	1.5048	1.49	1.51	–	1.51	1.599	
1.466	1.460	1.47	1.469	1.4714	1.47	1.48	1.471	1.46	1.485	

^aInternal position standard employed: (1) Fh75, NIST SRM640b, d-spacings in parentheses are tentative due to presence as broad reflections with low intensity; (2) BC-11, halite (NaCl) PDF 05-0628.

^bPrepared by hydrolysis of FeSO₄ (0.1 mol L⁻¹) and Na₂SiO₃ (ca. 0.015 mol L⁻¹) at pH 7 and room temperature.

^cReflections are also reported at 1.271 Å, 1.223 Å, and 1.104 Å; d-spacings at 4.61 and 1.104 Å are not indexed in PDF 13-0087. Also see *Drits et al.* [1993].

^d*Jambor et al.* [1995].

^ePeak identified through profile fitting; not clearly evident as separate peak in pattern.

^fNote that there is evidence of a shoulder on the high-angle side of the peak identified at 2.52 Å by *Janney et al.* [2000, Figure 1].

^g*Janney et al.* [2001] present a simulated diffraction pattern (their Figure 7) for 6-line ferrihydrite using a double-hexagonal structure with a reflection indexed as (101) that corresponds to the location of the peak observed at ~2.45 Å for samples Fh75 and BC-11; also indexed as (101) by *Drits et al.* [1993].

Dormann et al. [1997]). When the magnetizations of individual superparamagnetic particles are “frozen” by cooling without an applied field, the magnetic moments are randomly oriented and the resulting induced magnetization is zero at 0 K. With increasing temperature and in the presence of an applied field, the magnetic moments begin to fluctuate and then align in the direction of the field, leading to an increase of the total magnetization [*Chantrell et al.*, 1991; *Dormann et al.*, 1997]. Above the blocking temperature (T_B), the temperature associated with the peak observed in the ZFC curves, the thermal energy of disorder is higher than the ordering energy associated with the applied field and anisotropy energy barrier ($K_{eff}V$), and a superparamagnetic behavior results, leading to a decrease of the total magnetization with increasing temperature following the Curie-Weiss law. The presence of a maximum in the ZFC susceptibility curve is therefore associated with the transition between magnet-

ically relaxing superparamagnetic (unblocked) and thermally stable magnetization (blocked state; at this point we are not discussing Néel temperature) [*Hansen and Mørup*, 1999]. Each particle is associated with a given volume and corresponding value of T_B . A wide or narrow distribution of particle volumes will produce a T_B distribution and will also result in a broad or narrow peak for the ZFC curve. As can be seen in Table 2 and Figure 3, the blocking temperature (T_B^{Main}) drops substantially with decreasing synthesis temperature, from above 300 K for sample *Fh75* to about 33 K for sample *Fh7.5*. For samples *Fh75* and *Fh50* there seems to be present an additional distribution of fine particles (T_B^{Fine}) with variable T_B s around 25 K. This population of extremely small particles is represented by a small shoulder on the ZFC curve; out of phase ac susceptibility in Figure 5 inset (see below) also present evidences of presence of these small particles with similar T_B^{Fine} . The ZFC/FC curves of dc

Table 2. Hysteresis Parameters Estimate at 2 K, T_B Determined at ZFC Curves and Particle Diameter Calculated by XRD (d_{XRD}) and by Using Blocking Temperature (d_{TB})^a

Sample (Number of Replicates)	Synthesis Temp., °C (Time, hours)	Composition Fe, wt% (Si, wt%)	M_S , Am ² /kg	H_C , T	T_B , K	d_{TB} , nm	d_{XRD} , nm
Fh7.5 (2)	7.5 (56.2)	51.7 (3.5) ^b	5.6	0.36	33	5.1	1.9(0.7)
Fh22 (5)	22 (2.8–25.6)	57.6 (1.9)	5.3	0.34	50	5.9	5.4(1.5)
Fh50 (5)	50 (0.53–1.95)	64.4 (1.3)	7.0	0.31	150	8.5	10.4(1.6)
Fh75 (5)	75 (0.15–0.47)	66.5 (1.7)	7.2	0.27	330	11.1	23.6(3.0)
BC-11	87–89 ^c	~40 (~10)	7.9	0.10	33, >400	5.1, >11.8	9.1(2.0)

^aOne standard deviation is shown in parentheses for the mean d_{XRD} calculated using four or five diffraction peaks and the Scherrer relationship. Except for synthesis temperature, reported data are for composited samples.

^bEstimated 4.4 wt% Si based on differences in dissolved Fe and Si at the beginning and end of the experiments relative to an assumed formula for ferrihydrite (Fe₅HO₈ · 4H₂O, 480 g/mole).

^cRate of ferrous iron oxidation and apparent time of ferrihydrite formation is unknown for this sample.

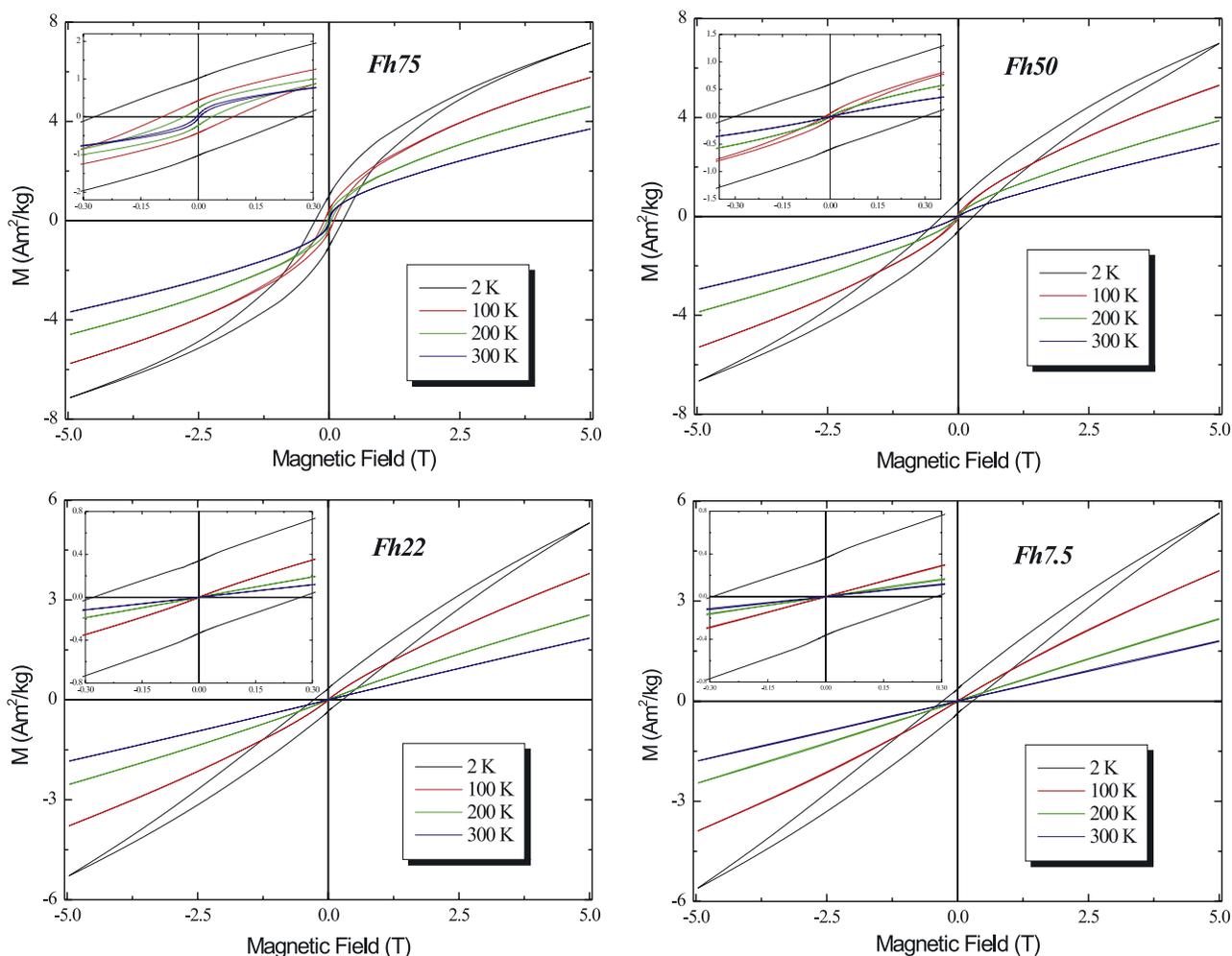


Figure 4. Hysteresis loops for *Fh75*, *Fh50*, *Fh22*, and *Fh7.5*. Data were collected at 2, 100, 200, and 300 K. Hysteresis loops did not attain saturation at 5 T.

susceptibility together with ac susceptibility information (see below) help us to fully infer the particle size distribution and its width.

[19] With the mean blocking temperature obtained from ZFC curve, it is possible to estimate the mean volume (V) of the particles through the relation:

$$V = \frac{25k_B T_B}{K},$$

where k_B is Boltzmann constant, K is the anisotropy constant, and T_B is the mean blocking temperature [Cullity, 1972; Chantrell et al., 1991]. By considering a spherical particle shape it is possible to obtain an estimate of the mean particle diameter. The expression above is very useful to estimate grain size from materials when the anisotropy constant is well determined. However, in the case of ferrihydrite, there is very little information on anisotropy constants in the literature. We have, however, used the magnetocrystalline anisotropy constant determined for typical particle size of ferrihydrite (~ 5 nm) of 1.7×10^5 J/m³ [Rodmacq, 1984]. Thus we could estimate the magnetic particle sizes for all four samples (see Table 2). The

diameters estimated by using the anisotropy constant above present a trend of increase with synthesis temperature like that observed by XRD. For the larger sample *Fh75* the value obtained (11.1 nm) is quite different from the one estimated by XRD (23 nm). The particle size disagreement observed in large particles could be associated with an incorrect value for the anisotropy constant used in our calculations. The anisotropy constant used in our estimate for all samples was the one obtained from typical ferrihydrite particle size (around 5 nm) where finite size effects, like surface anisotropy and dipolar interactions are important contributions, are taken into account. Since XRD data shows that this ferrihydrite sample (*Fh75*) clearly contains much larger grains (above 10 nm) the anisotropy constant used could be an overestimate, since the contributions from finite size are not so important for these larger particles. For sample *Fh7.5* effects due to interactions and/or the core/shell relationship could be related to the overestimate of particle size. Particles size estimates for samples *Fh22* and *Fh50* are in agreement considering the experimental error.

[20] The hysteresis loops at 2 K of all four synthetic samples are presented in Figure 4. Each sample displays a lack of saturation even at a magnetic field of 5 T. Another

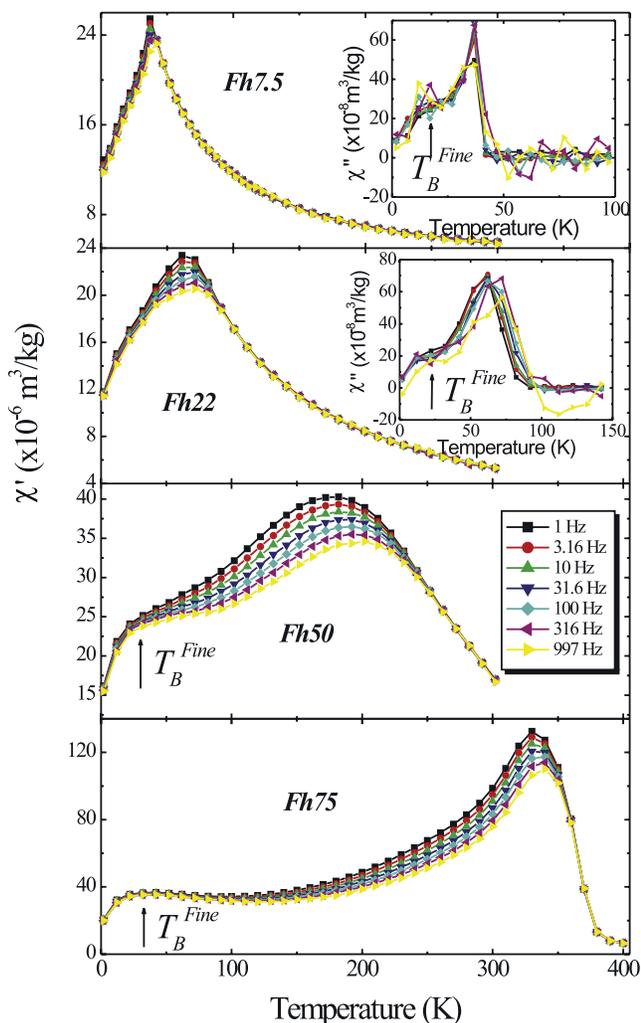


Figure 5. The ac susceptibility data for all for samples showing the peak temperature at different frequencies and the distribution of particles size with synthesis temperature. T_B^{Fine} represents the blocking temperature for the small size fraction.

interesting feature observed relates to coercive force (H_C) and remanent magnetization (M_R). The H_C values are similar but still show a trend of increase with increasing synthesis temperature, while the M_R values decrease with decreasing synthesis temperature (or decreasing particle size), although a small deviation from this trend is observed for sample *Fh22*. These two parameters are sensitive to grain size and composition variations and the M_R decrease with particle size could be related to Si-content. In our case, the variation of grain size is conspicuous, but we can not discard possible structural effects due to the incorporation of Si. The hysteresis loops obtained at 100 K, 200 K, and 300 K for samples *Fh22* and *Fh7.5* are represented mainly by a straight line with a small sigmoidal behavior, such as is expected from a superparamagnetic component; a similar behavior was found by *Zergenyi et al.* [1999] and *Guyodo et al.* [2006]. For the other two samples, *Fh75* and *Fh50*, a sigmoidal M versus B pattern was observed, and even an open hysteresis loop was obtained for sample *Fh75* at 300K.

These behaviors are in agreement with ZFC/FC curves, since above T_B we expect a pure superparamagnetic behavior very much like a paramagnetic material.

5.2. The ac Susceptibility

[21] Figure 5 displays the ac susceptibility $\chi(T, f)$ data for the four Si-ferrihydrite synthetic samples. Both real $\chi'(T)$ and imaginary $\chi''(T)$ components exhibit the expected behavior of a superparamagnetic system, that is to say, the occurrence of a maximum at a temperature T_m . At temperatures above T_m , χ' is frequency independent indicating thermodynamic equilibrium, as expected in a temperature range where all the particles are in a superparamagnetic state. However, at $T \leq T_m$, χ' is frequency-dependent, indicating a nonequilibrium situation, as expected when a fraction of particles are in the blocked state [Dormann *et al.*, 1997]. Since T_m depends on the measurement frequency, the peak occurs at different temperatures for different frequencies. The general behavior observed is the shift of this maximum toward higher temperatures with increase of the driving frequency. This information further supports the ZFC dc susceptibility data where we also observed the increase of blocking temperature (which represents the particle volume) as a consequence of the increase of synthesis temperature.

[22] The real component (χ') presents an additional peak around 25 K (related to very small particles) for samples *Fh75* and *Fh50* such as what was observed for the ZFC curves. From the imaginary component (χ'') we also could observe the contribution related to those small particles with $T_m \sim 25$ K for samples *Fh22* and *Fh7.5* (Figure 5 inset).

5.3. Mössbauer Spectroscopy

[23] Figure 6 shows the Mössbauer spectra at room temperature and 4.2 K for samples *Fh75* (largest) and *Fh7.5* (smallest). At 4.2 K the spectra for all four samples are only sextets, indicating a complete magnetic order. The hyperfine parameters, magnetic hyperfine field (B_{HF}), quadrupole splitting (QS), and isomer shift (IS), are presented in Table 3 and are in agreement with literature values of ferrihydrite [Murad and Cashion, 2004], although a small decrease of B_{HF} is seen when compared with a pure, “well” crystallized ferrihydrite. This decrease of B_{HF} could be associated with silicon substitution since it has been known that isomorphic substitution (doping) reduces the magnetic hyperfine field. On the other hand, we also know that lattice vacancies can produce a similar effect of decrease in B_{HF} . *Schwertmann et al.* [2004] reported a similar B_{HF} decrease due to incorporated Si. Sample *Fh7.5* has the maximum deviation when compared with the standard B_{HF} values of ferrihydrite at 4.2 K (50.8 T [Murad and Cashion, 2004]) and the X-ray diffraction patterns match with typical, poorly crystallized pattern of ferrihydrite. It seems there are two concurrent effects of the decrease in synthesis temperature: (1) an increase in the structural substitution of iron by Si and (2) a decrease in the crystallinity of the material (vacancies increase). The Mössbauer spectra at room temperature are represented by doublets for samples *Fh50*, *Fh22*, and *Fh7.5*. The exception is sample *Fh75*, which presents a collapsed magnetic hyperfine field, suggesting relaxation effects. Such an effect is in agreement with the magnetic data, which show an open hysteresis loop at room temperature

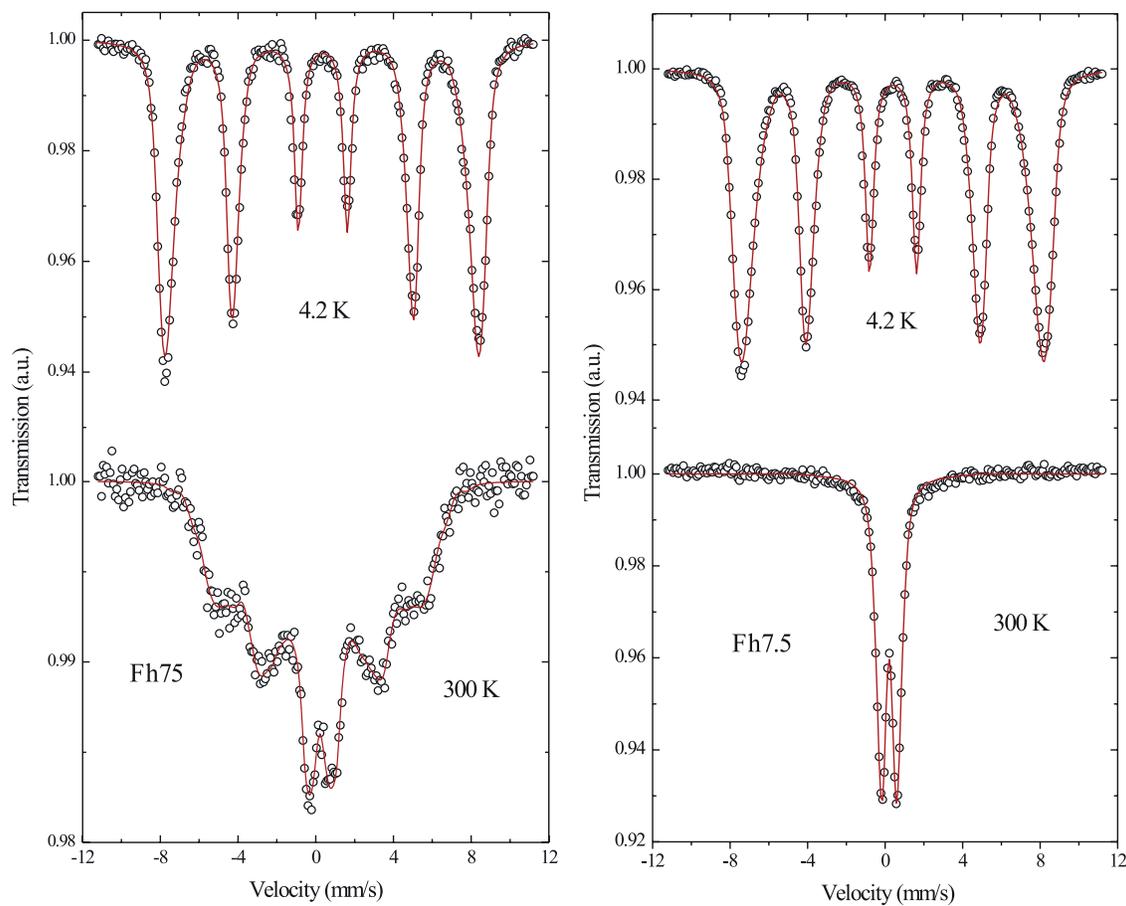


Figure 6. Mössbauer spectra at room temperature and 4.2 K for samples *Fh75* and *Fh7.5*. At 4.2 K the particles are blocked and the sextet corresponding to ferrihydrite was fitted. At room temperature *Fh7.5* present the usual superparamagnetic behavior represented by a doublet and *Fh75* shows a collapsed magnetic hyperfine field resulting from relaxation effects. Open circles are experimental points and line is the best fitting.

and a T_B^{Main} from ZFC curves and T_m from ac susceptibility above 300 K.

6. Particle Size Analysis of the Natural Ferrihydrite Particles (*BC-11*): Magnetic Information

[24] The magnetic information from sample *BC-11* seems to represent the sum of three distinct magnetic behaviors. The first is related to the presence of superparamagnetic particles

represented by the peak around 33 K in ZFC curve and in ac susceptibility curves (Figures 7a and 7b). The second is a paramagnetic contribution, which can be observed in the ZFC and ac susceptibility curves at temperatures below 10 K. Such a contribution is clear from the susceptibility increase at very low temperatures (2 K) and is confirmed by the fact no out of phase susceptibility component is observed (Figure 7b inset). The third is related to the remanent magnetization at room temperature. The frequency dependency in ac susceptibility curves clearly demonstrates a broad particle size

Table 3. Mössbauer Hyperfine Parameters at Room Temperature and 4.2 K for Synthetic and Natural Ferrihydrite^a

Temperature		Sample						
		Fh7.5	Fh22	Fh50	Fh75	BC-11	2-Line ^b	6-Line ^b
300 K	B_{HF} , T	-	-	-	33.3(1)	-	-	-
	QS, mm/s	0.78(1)	0.75(1)	0.72(1)	-0.07(3)	0.78(1)	0.80	0.62
	IS, mm/s	0.34(1)	0.36(1)	0.35(1)	0.37(1)	0.34(1)	0.35	0.35
4.2 K	B_{HF} , T	47.8(5)	48.5(1)	49.4(1)	49.7(2)	49.1(1)	47	50
	QS, mm/s	-0.02(1)	-0.03(1)	-0.04(1)	-0.04(1)	-0.04(1)	-0.02	-0.07
	IS, mm/s	0.43(1)	0.46(1)	0.46(1)	0.47(1)	0.55(1)		

^aMaximum hyperfine field (B_{HF}), quadrupole splitting (QS), isomer shift (IS). Errors are quoted within parentheses.

^bMurad and Cashion [2004].

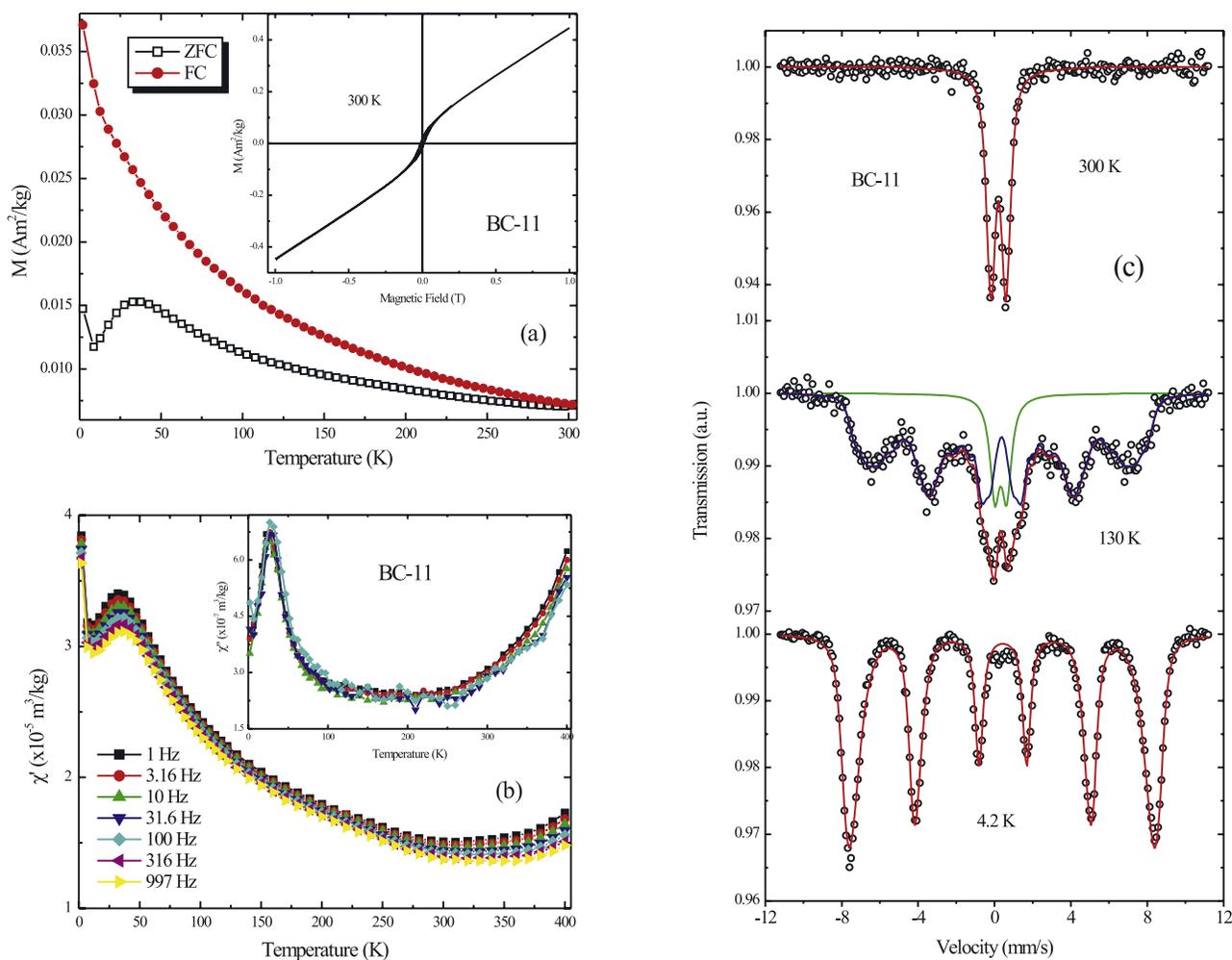


Figure 7. (a) ZFC/FC curves with applied field of 5 mT, inset: hysteresis loops at room temperature; (b) ac susceptibility in phase, inset: ac susceptibility out phase; (c) Mössbauer spectra at room temperature, 130 K and 4.2 K from natural ferrihydrite sample (*BC-11*).

distribution. Furthermore, the out of phase susceptibility (Figure 7b inset) increases from 300 to 400 K and beyond. This behavior is consistent with the presence of a large particle size fraction. It is the larger particles that are responsible for the lack of overlap between ZFC and FC curves. Interestingly, *BC-11* and *Fh75* share an important magnetic feature, hysteresis at room temperature (Figure 7a inset). This is strong evidence for behavior outside of the exclusively superparamagnetic threshold of the vast majority of ferrihydrite heretofore characterized. This is to say, the open hysteresis loops (displaying finite coercive force) are therefore related to the presence of a fraction of larger particles.

[25] Finally, Mössbauer spectra (Figure 7c) collected for *BC-11* demonstrate thermally blocked magnetization at 4.2 K, at 130 K two components corresponding to a doublet and sextet were fitted and at room temperature only a doublet, which signifies the superparamagnetic threshold. The hyperfine parameters fitted at 4.2 K corresponding to the sextet are in agreement with ferrihydrite and are close to the parameters found for sample *Fh75* (particle size = 23 nm). Interestingly, complete magnetic order (i.e., a

sextet), which should be associated with the large particles at room temperature (and evident in all other types of magnetic data) is not present in the Mössbauer spectrum at 300 K. Such a sextet was shown by the synthetic sample *Fh75* at 300 K. This feature could indicate that the fraction of large ferrihydrite is simply below the detection limit of Mössbauer spectroscopy. Another explanation worth considering is related to degree of agglomeration of natural ferrihydrite. *Mørup et al.* [2002] have presented evidence on the importance of agglomeration for nanophase hematite. In their work, they present Mössbauer spectra evolving from 4.2 K to room temperature of the same particle size with coated and uncoated surfaces. Interestingly, the Mössbauer spectra of *BC-11* follow the same pattern as coated particles. We present the spectrum at 130 K (Figure 7c) where the coexistence of a doublet and sextet is noticed instead of the collapsed magnetic hyperfine field typical of interacting particles. *Mørup et al.* [2002] show (in their Figure 1) that the Mössbauer spectra at room temperature have different patterns in the two cases, a doublet for coated and a sextet for uncoated samples. Also, they noted that a substantial decrease of temperature was necessary for onset of magnetic

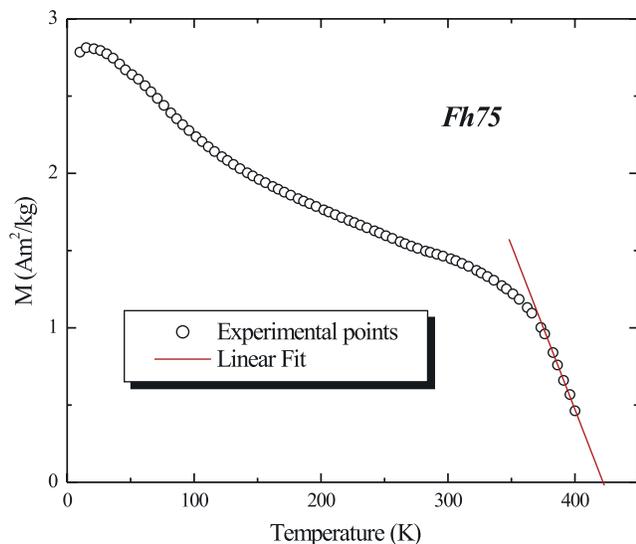


Figure 8. Induced magnetization curve (measure with 1 T magnetic field applied) showing the decrease of magnetization with temperature. At 400 K the sample carried only 17% of the initial induced magnetization. $T_N = 422$ K was obtained by extrapolation.

order for the coated sample when the doublet was the main feature. HRTEM of *BC-11* ferrihydrite (Figure 2b) demonstrates that the particles are clearly in contact with one another. However, Mössbauer results suggest that the ferrihydrite nanoparticles are noninteracting. This interpretation is supported by the sextet and doublet superposition as observed for noninteracting particles. Thus we propose that the natural ferrihydrite nanoparticles have a core-shell structure in which the shell is enriched in Si/As, which could prevent magnetic interactions due to aggregation. However, a small fraction of the particles could be agglomerated, and in the presence of a large steady magnetic field (1 T) they behave like larger particles and are also responsible for a finite coercive force displayed at 300 K.

7. Can We Determine Ferrihydrite Néel Temperature?

[26] Since it was observed that T_B^{Main} is above room temperature for sample *Fh75* ($d_{XRD} = 23.6$ nm) our natural inclination was to attempt to determine Néel temperature (T_N), a very useful and intrinsic property to identify magnetic phases. For ferrihydrite, T_N is not well known since relaxation effects due to small particle size are observed at low temperatures (in the range of 60 K and 150 K). The usual superparamagnetic behavior could thus be misinterpreted as a paramagnetic state. The main source of information about ferrihydrite T_N comes from fitting magnetization curves of ferritin using the Langevin model [Gilles *et al.*, 2000]. The extrapolated temperatures obtained by other authors [e.g., Guyodo *et al.*, 2006] claim the T_N of synthetic ferrihydrite to be between 400 K and 500 K. To the best of our knowledge, there is no experimental evidence obtained with larger size ferrihydrite where the behavior is not superparamagnetic but thermally stable at room temperature. In order to determine T_N we cooled the 23 nm sample (*Fh75*) from room temper-

ature down to 5 K and then applied a magnetic field of 1 T to suppress effects of superparamagnetic particles. The induced magnetization was then measured with increasing temperature up to 400 K. The data are presented in Figure 8. We can see the slow magnetization decrease, and at our maximum experimental temperature of 400 K, approximately 17% of the magnetization is still left, suggesting that T_N is between 400 and 500 K for *Fh75*. Fitting a straight line with a constant slope for the final part of the curve provides an extrapolated T_N of 422 K for our Si-ferrihydrite. This temperature is in agreement within the temperature range estimated by Guyodo *et al.* [2006] although our T_N may be slightly smaller due to Si-doping effects. It is worth remembering that the samples studied by Guyodo *et al.* [2006] were pure ferrihydrite. A similar effect of decrease of Curie temperature in a doped material is observed in other iron oxide phases like titanomagnetite [Cullity, 1972; O'Reilly, 1984].

8. Conclusion

[27] The magnetic characterization performed with these silicon-bearing ferrihydrites showed an important feature: ferrihydrite particle diameter depends on the synthesis temperature. Magnetic characterization, in combination with X-ray diffraction and HRTEM, demonstrates that the size of ferrihydrite particles produced by air oxidation of aqueous ferrous solution increases with increasing synthesis temperature. Sample *BC-11* collected at a marine shallow-water hydrothermal system in Bahia Concepcion, Baja California, Mexico is an example of large natural ferrihydrite. The hydrothermal fluid temperature is around 88°C and the conditions (high temperatures and Si concentration) seem to favor the growth of such large ferrihydrite. Although large particle size may not be common in nature, the sample *BC-11* presented a “7-line” XRD pattern as well as a signature for stable single domain particles, easily identified by open hysteresis loops. For all five samples, (synthetic and natural ferrihydrite) an important contribution of small particles with T_B^{Fine} around 25 K was observed. Those small particles could reflect the presence of primary particles (or seeds) that may be related to the formation of the larger ferrihydrite particles. A synthetic sample (*Fh 75*) with large particle size, the seven-line silicon-doped ferrihydrite described here, yields an estimated T_N of 422 K. Our results suggest that very large crystallites of ferrihydrite could act as carriers of natural remanent magnetization, but it is improbable that particles smaller than 10 nm could. The rare occurrence of large particles means that ferrihydrite will not likely contribute to magnetic record. Other topics that need to be investigated in the future include ferrihydrite anisotropy constant since, for this material, there is not a reliable value available as for other iron phases.

[28] **Acknowledgments.** This study was supported by National Science Foundation (NSF) grant EAR 0311869 from the Biogeosciences program and MRI EAR 0320641. Institute for Rock Magnetism (IRM) is funded by the Earth Science Division of NSF, the W. M. Keck Foundation and University of Minnesota. This is IRM publication 0508. The use of trade names does not imply endorsement. The U. S. Environmental Protection Agency through its Office of Research and Development funded a portion of the research described herein. It has not been subjected to Agency review and therefore does not necessarily reflect the views of the Agency, and no official endorsement should be inferred.

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